Name _____ Date_____

It's a Balancing Act!!

Examine the equation. Decide if it has an equal number of each type of atom on each side of the arrow. Try to add coefficients to balance if it is not in equilibrium.
Atoms of REACTANTS = Atoms of PRODUCTS

Equation	Is it balanced?
$N_2 + H_2 \rightarrow NH_3$	
$KClO_3 \rightarrow KCl + O_2$	
$NaCl + F_2 \rightarrow NaF + Cl_2$	
$AgNO_3 + MgCl_2 \rightarrow AgCl + Mg(NO_3)_2$	
$CH_4 + O_2 \rightarrow CO_2 + H_2O$	
$S_8 + O_2 \rightarrow SO_3$	
$P + O_2 \rightarrow P_2O_5$	
$K + MgBr_2 \rightarrow KBr + Mg$	
$HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$	
$Na + Cl_2 \rightarrow 2NaCl$	
$H_2 + O_2 \rightarrow H_2O$	
$SnO_2 + 2C \rightarrow 2CO + Sn$	
$Ca + Cl_2 \rightarrow CaCl_2$	
$C + H_2 \rightarrow CH_4$	
$Zn + H_2SO_4 \rightarrow Zn SO_4 + H_2$	
$NO + O_2 \rightarrow NO_2$	
$NaOH + HCl \rightarrow NaCl + H_2O$	
$CdCO_3 \rightarrow CdO + CO_2$	
$KI + H_2O_2 \rightarrow KOH + I_2$	
$Ag + O_2 \rightarrow Ag_2O_2$	
$HgCl_2 + 2NH_3 \rightarrow NH_4Cl + HgNH_2Cl$	