

**Law of Conservation of Mass Worksheet**

Name \_\_\_\_\_ Per. \_\_\_\_\_

The law of conservation of mass says:

*Fill in the blank with the correct number.*

1. 40 g of calcium reacts with 71 g of chlorine to produce \_\_\_\_\_ g of calcium chloride.

2. \_\_\_\_\_ g of potassium reacts with 16 g of oxygen to produce 94 g of potassium oxide.

*Example:*  $K + 16 = 94$

$K = ?$

3. 14 g of lithium reaction with \_\_\_\_\_ g sulfur to produce 46 g of lithium sulfide.

4. 24 g of magnesium reacts with 38 g of fluorine to produce \_\_\_\_\_ g magnesium fluoride.

5. 65.5 g copper reacts with \_\_\_\_\_ g oxygen to produce 81 g copper (I) oxide.

**Law of Conservation of Mass Worksheet**

Name \_\_\_\_\_ Per. \_\_\_\_\_

The law of conservation of mass says:

*Fill in the blanks with the correct number.*

1. 88 g of strontium reacts with 160 g bromine to produce \_\_\_\_\_ g strontium bromide.

2. 46 g of sodium reacts with \_\_\_\_\_ g chlorine to produce 117 g sodium chloride.

*Example:*  $46 + \text{Cl} = 117$

$\text{Cl} = ? \text{ g}$

3. \_\_\_\_\_ g iron reacts with 71 g chlorine to produce 129 g of iron (II) chloride.

4. 137 g of barium reacts with \_\_\_\_\_ g iodine to produce 391 g barium iodide.

5. \_\_\_\_\_ g hydrogen reacts with 32 g of oxygen to produce 34 g of hydrogen peroxide.