Name:			<i>Period:</i>
Physical versus Chemical Pr	operties		
I. Reviewing matter:	•		
Matter: anything that has	_ and takes up		
<u>Mass</u> – the amount of in sor	nething		
<u>Volume</u> – the amount ofs	omething occup	ies	
Is it Matter?	Yes	No	
A car?	•	•	
A box?			
You?			
Heat?			
II. Property: a characteristic of a			
III. Physical property: a property that			
the substance. Examples: luster,			
), (the ab), melting point,
point, density, solubility	y and specific he	at.	
IV. Special properties:	which a cubetan	aa ahangaa fram a	solid to a stagiyon
Melting point : temperature at which the temperature at		_	solid to a at a given
Boiling point: temperature at v			to a at a
given pressure.			00 ti 40 ti
H ₂ O = or			
V. Chemical property: a property that		observed by	the
of the substance. Examples: flammabili	ty, ability to rus	t, reactivity with v	rinegar
VI. Density: the amount of		`	
Density can be used to identify a substa			
Water's density is			
VII. Calculations $D = m/V = g/mL = g$	/cm3		
Examples: A cube has a mass if 2.8g	and occupies a	volume of 3 67mI	. Would this object float or
sink in water? This object would	-		•
		ater because its ut	ensity is than
water whose density is	-•		
VIII. More Density Calculations			
A liquid has a mass of 25.6 g and a v	olume of 31.6 n	ıL.	
What is the identity of the liquid? _			
*Use the information in the chart fo	r reference.		
Substance			Density (g/ml)

Substance	Density (g/mL)
Mercury	13.6
Water	1.0
Ethanol	0.81

Chemical versus Physical Properties

Property Description	Chemical	Physical
Can react with vinegar		
Density		
Can react with the oxygen in the air		
Luster (shininess)		
The ability to freeze		
Can react with an acid		
Combustible		
The ability to melt		
The ability to digest food		
The ability to sublime (solid →gas)		
Malleability		
Ductility		
The ability to react with water		
The ability to neutralize stomach acid		
Color		
Magnetism		
Odor		
The ability to rust		
The ability to evaporate		